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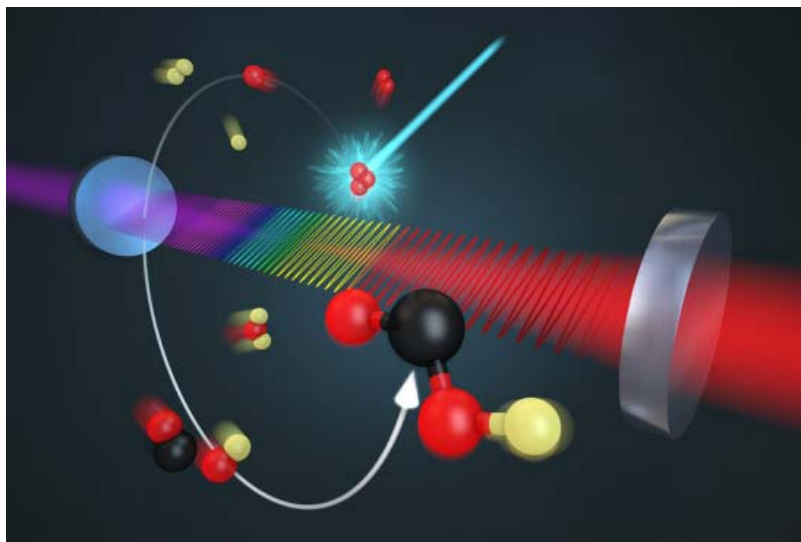
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JILA team identifies missing piece in how fossil fuel combustion contributes to air pollution

31 October 2016

JILA physicists and colleagues have [identified](#) a long-missing piece in the puzzle of exactly how fossil fuel combustion contributes to air pollution and a warming climate. Performing chemistry experiments in a new way, they observed a key molecule that appears briefly during a common chemical reaction in the atmosphere. JILA a partnership of the National Institute of Standards and Technology (NIST) and the University of Colorado Boulder. Their paper appears in *Science*.

The reaction combines the hydroxyl molecule (OH, produced by reaction of oxygen and water) and carbon monoxide (CO, a byproduct of incomplete fossil fuel combustion) to form hydrogen (H) and carbon dioxide (CO₂).



Artist's conception of an infrared frequency comb "watching" the reaction of a molecule of carbon monoxide (CO, red and black) and hydroxyl radical (OH, red and yellow) as they form the elusive reaction intermediate DOCO (red/black/red/yellow) before DOCO falls apart. This chemical reaction was seen for the first time under normal atmospheric conditions in the laboratory by the Ye group. Credit: Jun Ye group and Steve Burrows/JILA Click to enlarge.

Researchers have been studying this reaction for decades and observed that its speed has an abnormal pressure and temperature dependence, suggesting there is a short-lived intermediate, the hydrocarboxyl molecule, or HOCHO. But until now, HOCHO had not been observed directly under conditions like those in nature, so researchers were unable to calculate accurately the pressures at which the reaction either pauses at the HOCHO stage or proceeds rapidly to create the final products.

As described in their paper, JILA's direct detection of the intermediate compound and measurements of its rise and fall under different pressures and different mixtures of atmospheric gases revealed the reaction mechanism, quantified product yields, and tested theoretical models that were incomplete despite rigorous efforts.

We followed the reaction step by step in time, including seeing the short-lived, and thus elusive, intermediates that play decisive roles in the final products. By finally understanding the reaction in full, we can model the atmospheric chemical processes much more accurately, including how air pollution forms.

—JILA/NIST Fellow Jun Ye

JILA researchers used a laser to induce the reaction inside a container called a laboratory flow cell, through which samples of the molecules participating in the reaction and other gases passed. This process mimicked nature by using gases found in the atmosphere and no catalysts.

To avoid any confusion in the results due to the presence of water (which contains hydrogen), the researchers used deuterium, or heavy hydrogen, in the hydroxyl molecule, OD, to start the reaction. Thus, they looked for the DOCO intermediate instead of HOCO. During the experiment, concentrations of CO and nitrogen gases were varied across a range of pressures.

Using JILA's patented frequency comb spectroscopy technique, which identifies chemicals and measures their concentrations in real time based on colors of light they absorb, researchers measured the initial OD and the resulting DOCO over various pressures and atmospheric gas concentrations over time, looking for conditions under which DOCO stabilized or decomposed to form CO₂.

The JILA team identified an important factor to be energy transfer due to collisions between the intermediate molecule and nearby CO and nitrogen molecules. These collisions can either stabilize the intermediate DOCO or deactivate it and encourage the reaction to proceed to its final products. JILA's frequency comb spectroscopy technique analyzes chemicals inside a glass container, in which comb light bounces back and forth between two mirrors. The repeated, continuous measurements make the technique especially sensitive and accurate in identifying "fingerprints" of specific molecules. This latest experiment used new "supermirrors," which have crystalline coatings that reduce light losses and improved detection sensitivity 10-fold.

JILA's results, notably the effects of molecular collisions, need to be included in future atmospheric and combustion model predictions, according to the paper.

For example, even at low pressures, the reaction produces a DOCO yield of nearly 50%, meaning about half the reactions pause at the intermediate stage. This observation affects calculations that go beyond Earth: Other researchers have shown that HOCO can contribute 25-70% of the total CO₂ concentration in the cold Martian atmosphere.

In the future, JILA researchers plan to extend the experimental approach to study other chemical products and processes. One topic of interest is reactions involving water and CO₂, to aid understanding of how atmospheric CO₂ interacts with and acidifies the oceans. Also of interest are studies of engine combustion, which affects fuel economy. A car engine combines air (oxygen and nitrogen) and fuel (hydrocarbons) to produce CO₂ and water. Incomplete combustion creates CO.

A physical chemist at the California Institute of Technology collaborated on the research. Crystalline Mirror Solutions provided the mirrors. The research was funded by the Air Force Office of Scientific Research, Defense Advanced Research Projects Agency, NIST and the National Science Foundation.

Resources

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